Students are expected to keep up with class work when absent.

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<td>● Assignment 2.0&lt;br&gt;● Assignment 2.1&lt;br&gt;● Read section(s) 2.8</td>
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<td>● 2.8 – Naming Simple Compounds&lt;br&gt;○ Type I Binary&lt;br&gt;○ Type II Binary&lt;br&gt;○ Type III Binary</td>
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<td>3</td>
<td>● 2.8 – Naming Simple Compounds&lt;br&gt;○ Ternary Compounds&lt;br&gt;○ Acids that do NOT contain oxygen&lt;br&gt;○ Acids the DO contain oxygen</td>
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<td>● Grade &amp; discuss assignment</td>
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<td>5</td>
<td>● Chapter 2 Test</td>
<td>● Read section(s) 3.2 – 3.5</td>
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Quiz 2.5
2.5 – The Modern View of Atomic Structure
1. What determines the chemistry of an atom?
2. Regarding nuclear symbols:
   a. What is the top number called?
   b. What does the top number represent?
   c. What is the bottom number called?
   d. What does the bottom number represent?

Quiz 2.6
2.6 – Molecules & Ions
3. What name is given to the force that holds atoms together in a compound?
4. What happens with the electrons when a covalent bond is formed?
5. What happens with the electrons when an ionic bond is formed?
6. Know the difference between a molecule and an ion.
7. What name is given to an ion with a positive charge?
8. What name is given to an ion with a negative charge?

Quiz 2.7
2.7 – An Introduction to the Periodic Table
9. Know the locations of the following groups of elements on the periodic table.
   a. metals
   b. nonmetals
   c. alkali metals
   d. alkaline earth metals
   e. halogens
   f. noble gases
   g. transition metals

Quiz 2.8
2.8 – Naming Simple Compounds
10. What is meant by the term “binary compound”?
11. When naming a binary compound what comes first? second?
12. Be able to name and write formulas for simple, binary compounds.
13. Be able to name and write formulas for simple, ternary compounds.
14. Be able to name and write formulas for simple acids.
At the completion of chapter 2 you should…

1. Know the definitions of the following terms.
   a. Proton
   b. Neutron
   c. Electron
   d. Isotope
   e. Atomic number
   f. Mass number

2. Understand what differentiates isotopes of a given element.
   a. Be able to write the nuclear symbol for a given isotope
   b. Be able to tell the number of protons, neutrons, and electrons in a given isotope

3. Understand what differentiates molecules and ions

4. Understand the basics of the periodic table
   a. Where are the metals found?
   b. Where are the nonmetals found?
   c. Where are the metalloids found?
   d. What is meant by the term Group?
   e. What is meant by the term Period?

5. Be able to name and write chemical formulas for
   a. Type I binary compounds
   b. Type II binary compounds
   c. Type III binary compounds
      i. Using the “old system” (prefixes)
      ii. Using the “new system” (Roman numerals)
   d. Ternary compounds
      i. Also includes naming system for polyatomic ions that contain oxygen.
   e. Acids
      i. Acids that do NOT contain oxygen (binary acids)
      ii. Acids that DO contain oxygen (ternary acids)
Assignment 2.0 – Vocabulary

Define each of the following terms.

1) Proton

2) Neutron

3) Electron

4) Isotope

5) Atomic number

6) Mass number
Assignment 2.1

1) Distinguish between the following terms:
   A) molecule versus ion
   B) covalent bonding versus ionic bonding
   C) molecule versus compound
   D) anion versus cation

2) The number of protons in an atom determines the identity of the atom.
   A) What does the number and arrangement of electrons in an atom determine?
   B) What does the number of neutrons in an atom determine?

3) Indicate the number of protons, neutrons, and electrons in each of the following atoms.
   A) $^{235}_{92}U$
   B) $^{13}_{6}C$
   C) $^{57}_{26}Fe$
   D) $^{208}_{82}Pb$
   E) $^{86}_{37}Rb$
   F) $^{41}_{20}Ca$
4) What is the trend in metallic character going from left to right across a period in the periodic table?

5) Consider the elements in the carbon family. What is the trend in metallic character going down this group?
Assignment 2.2 – Nomenclature

1) Provide the correct chemical name for each of the following:

A) NaBr

B) Rb₂O

C) CaS

D) AlI₃

E) Hg₂O

F) FeBr₃

G) CoS

H) TiCl₄

I) Na₂CO₃

J) NaHCO₃

K) S₄N₄

L) BaCrO₄

M) NH₄NO₃

N) N₂O

O) SO₃

P) Cl₂F

Q) HBr

R) HI

S) HNO₂

T) H₂SO₄
2) Give the correct chemical formula for each of the following:

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<tbody>
<tr>
<td>A</td>
<td>strontium fluoride</td>
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<tr>
<td>B</td>
<td>aluminum selenide</td>
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<tr>
<td>C</td>
<td>potassium nitride</td>
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<td>D</td>
<td>magnesium phosphide</td>
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<tr>
<td>E</td>
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